
CHEMISTRY

9701/22

Paper 2 AS Structured Questions

October/November 2019

MARK SCHEME

Maximum Mark: 60

Published

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes should be read in conjunction with the question paper and the Principal Examiner Report for Teachers.

Cambridge International will not enter into discussions about these mark schemes.

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This document consists of **10** printed pages.

PUBLISHED**Generic Marking Principles**

These general marking principles must be applied by all examiners when marking candidate answers. They should be applied alongside the specific content of the mark scheme or generic level descriptors for a question. Each question paper and mark scheme will also comply with these marking principles.

GENERIC MARKING PRINCIPLE 1:

Marks must be awarded in line with:

- the specific content of the mark scheme or the generic level descriptors for the question
- the specific skills defined in the mark scheme or in the generic level descriptors for the question
- the standard of response required by a candidate as exemplified by the standardisation scripts.

GENERIC MARKING PRINCIPLE 2:

Marks awarded are always **whole marks** (not half marks, or other fractions).

GENERIC MARKING PRINCIPLE 3:

Marks must be awarded **positively**:

- marks are awarded for correct/valid answers, as defined in the mark scheme. However, credit is given for valid answers which go beyond the scope of the syllabus and mark scheme, referring to your Team Leader as appropriate
- marks are awarded when candidates clearly demonstrate what they know and can do
- marks are not deducted for errors
- marks are not deducted for omissions
- answers should only be judged on the quality of spelling, punctuation and grammar when these features are specifically assessed by the question as indicated by the mark scheme. The meaning, however, should be unambiguous.

GENERIC MARKING PRINCIPLE 4:

Rules must be applied consistently e.g. in situations where candidates have not followed instructions or in the application of generic level descriptors.



GENERIC MARKING PRINCIPLE 5:

Marks should be awarded using the full range of marks defined in the mark scheme for the question (however; the use of the full mark range may be limited according to the quality of the candidate responses seen).

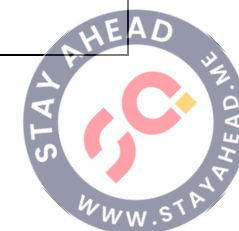
GENERIC MARKING PRINCIPLE 6:

Marks awarded are based solely on the requirements as defined in the mark scheme. Marks should not be awarded with grade thresholds or grade descriptors in mind.



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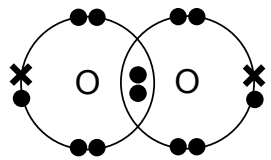
Question	Answer	Marks
1(a)(i)		1
1(a)(ii)	$\text{Si(g)} \rightarrow \text{Si}^+(\text{g}) + \text{e}^-$	1
1(a)(iii)	M1: similar shielding AND increase in proton number / atomic number / nuclear charge M2: increased nuclear attraction	2
1(a)(iv)	M1: 3 OR 13 M2: large(r) <i>increase</i> between third and fourth ionisation energies OR large(r) <i>increase</i> after third electron removed	2
1(b)(i)	M1: $\frac{(92.2 \times 28) + ({}^{29}\text{Si} \times 29) + ({}^{30}\text{Si} \times 30)}{100} = 28.09$ M2: ($x =$) 6.6 OR $28.09 = 28.078 + x$ (where x = abundance of Si-29) M3: 7.8 – <i>M2 calculated correctly to one decimal place (or more)</i> (=) 1.2%	3
1(b)(ii)	M1 giant (molecule) M2 strong covalent bonds (between atoms / particles) M3 no mobile charged particles / carriers	3
1(c)(i)	$\text{C}_2\text{H}_5\text{SH} + 4\frac{1}{2}\text{O}_2 \rightarrow 2\text{CO}_2 + 3\text{H}_2\text{O} + \text{SO}_2$	1

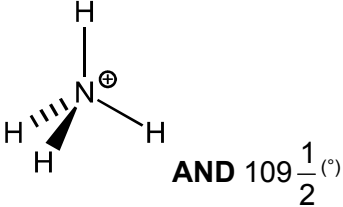


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Question	Answer	Marks
1(c)(ii)	M1: (causes) acid rain OR reacts/dissolves with (rain)water (vapour) to form (sulfuric / sulfurous) acid M2: one point from the following list. • lowers pH / increases acidity of rivers / lakes / oceans / water supplies / seas / soil / ground water • kills/harms / damages fish • kills / harms / damages plants / damages coral / aquatic life / plants / crops / trees or deforestation • leaches (toxic) aluminium (ions / salts) from soil (into rivers / lakes) • leaches away soil nutrients / soil unfit for agriculture • damages / weathers / erodes / destroys buildings / statues	2
1(d)(i)	M1: no effect / none M2: equal mol(es) (<i>of gas</i>) on both sides (of equilibrium / equation) owtte	2
1(d)(ii)	M1: (forward reaction is) endothermic M2: Any temperature higher than 300 K	2



Question	Answer	Marks										
2(a)	<table><tr><td>Na₂O</td><td>Al₂O₃</td><td>SiO₂</td><td>P₄O₁₀ / P₄O₆ / P₂O₃ / P₂O₅</td><td>SO₃</td></tr><tr><td>basic</td><td>amphoteric</td><td>acidic</td><td>acidic</td><td>acidic</td></tr></table> M1: all formulae correct M2: all acid / base behaviour correctly stated	Na ₂ O	Al ₂ O ₃	SiO ₂	P ₄ O ₁₀ / P ₄ O ₆ / P ₂ O ₃ / P ₂ O ₅	SO ₃	basic	amphoteric	acidic	acidic	acidic	2
Na ₂ O	Al ₂ O ₃	SiO ₂	P ₄ O ₁₀ / P ₄ O ₆ / P ₂ O ₃ / P ₂ O ₅	SO ₃								
basic	amphoteric	acidic	acidic	acidic								
2(b)(i)	reacts with both acid and base	1										
2(b)(ii)	OH ⁻ + H ⁺ → H ₂ O	1										
2(b)(iii)	reaction with strontium / reaction 1 will effervesce / fizz / bubble OR no fizzes / bubbles / effervescence with SrO	1										
2(b)(iv)	increases	1										
2(c)(i)	M1: correct conversion of quantities V = 5(.00) × 10⁻⁴ (m³) T = 293 (K) p = 5.37(0) × 10³ (Pa) M2: calculation to find n using n = PV / RT (n) = 1.1 × 10⁻³ (mol)	2										
2(c)(ii)	(i) × 78 = 0.0860 g	1										
2(c)(iii)	 M1: bonding pair between the two O M2: total of 14 electrons distributed equally between the two O	2										

Question	Answer	Marks
3(a)(i)	green gas fades OR white solid / white powder / white smoke / white fumes	1
3(a)(ii)	hydrolysis	1
3(a)(iii)	• P goes from 0 to (+)5 / (+)V • P is oxidised • N goes from (+)5 / (+)V to (+)4 / (+)IV • N is reduced Award one mark for two correct bullet points, award two marks for all four correct.	2
3(b)(i)	accepts a proton / H^+ OR donates a (lone) pair of e^-	1
3(b)(ii)	3-D shape AND bond angle  AND $109\frac{1}{2}^{\circ}$	1
3(b)(iii)	fertilisers	1
3(c)(i)	$C_2H_5OH + PCl_5 \rightarrow C_2H_5Cl + POCl_3 + HCl$	1
3(c)(ii)	substitution	1

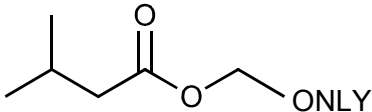
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Question	Answer	Marks
3(c)(iii)	EITHER M1: HI / I⁻ is a strong(er) reducing agent (than HCl / Cl⁻) M2: HI / I⁻ is oxidised (to iodine but the chloride is not) OR M1: H₂SO₄ is a (strong enough) oxidising agent (to react with HI / I⁻ here) M2: HI / I⁻ forms iodine OR M1: phosphoric acid is a weak / not an oxidising agent M2: (so) does not react with iodide (where M2 is dependent on M1 here)	2
3(c)(iv)	M1: C₂H₅I reacts fastest AND C₂H₅Cl reacts slowest OR C₂H₅Cl < C₂H₅Br < C₂H₅I M2: C—I bond is the weak(est) AND C—Cl bond strong(est)	2



Question	Answer	Marks
4(a)	M1: curly arrow from lone pair on O of OH⁻ to C of C—Cl M2: correct dipole on C—Cl AND curly arrow from C—Cl bond to Cl^{δ-}	2
4(b)(i)		1
4(b)(ii)	(CH ₃) ₂ CHCH(OH)CH ₂ OH	1
4(b)(iii)	optical (isomerism)	1
4(b)(iv)	C ₅ H ₁₂ O ₂ + 3[O] → C ₅ H ₈ O ₃ + 2H ₂ O	1
4(c)(i)	Add bromine water / Br ₂ (aq) AND turns (from orange / brown to) colourless	1
4(c)(ii)		1
4(d)(i)	3-methylbutan-1-ol	1
4(d)(ii)	heterogeneous	1

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Question	Answer	Marks
4(d)(iii)	M1: skeletal formula of Q  M2: one commercial use of Q (ethyl isovalerate / ethyl 3methylbutyrate) solvents / perfumes / flavourings	2
4(e)(i)	1500–1680 (cm ⁻¹) AND C=C	1
4(e)(ii)	potassium cyanide / KCN / sodium cyanide / NaCN	1
4(e)(iii)	(acidic) hydrolysis	1
4(e)(iv)	M1: recognise this reaction involves less stable intermediate 1° (carbo)cation (intermediate) is less stable (than 3°) M2: explain difference in reactivity in terms of positive inductive effect – comparative answer lower (positive) inductive effect / lower (+)I OR inductive effect of less alkyl groups	2

